Notes re Quiz 1 (Chem 6A, W'03)

Can't have 3 bonds to H.

- C\equiv H is not possible. It does not represent \( sp^2 \) C-H hybrids. \( sp^2 \) is a superscript (not a subscript).

\[ \text{C} = \text{C} = \text{C} \]

When "H" is written out, the atoms attached to it must be specified.

Similarly:

\[ \text{C} = \text{C} = \text{C} \]

is the same as \[ \text{C} \]

not the same as \[ \text{CH}_3 \]. When a line is drawn from carbon (to represent a bond), that does not specify an atom or group at the end of the line, it is always assumed that it terminates in a \( \text{CH}_3 \).

\[ \triangle \]

is not \[ \Delta \]. The H must be added.

\[ \text{OH} \rightarrow \text{O} \]  
I gave credit for the end form - this time. You will learn